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26 Kinetics, Spring 2008

Entropy and Second Law of
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Chemical Thermodynamics:

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Energy - Equilibrium

Constant, Enthalpy \u0026

Entropy - Equations \u0026

Practice Problems 19.1

Thermodynamics **chapter 18.**

free energy, spontaneity and

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equilibrium *Thermodynamics:*

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Thermodynamics

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Using the second law of

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thermodynamics, explain why heat flows from a hot body to a cold body but not from a cold body to a hot body. One test of the spontaneity of a reaction is whether the entropy of the universe increases: $\Delta S_{\text{univ}} > 0$.

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Using an entropic argument,
show that the following
reaction is spontaneous at
25°C:

19.E: Chemical
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chapter. Lecture Outline

19.1 Spontaneous

Processes^{1,2,3,4} • Chemical thermodynamics is concerned with energy relationships in chemical reactions. • We consider enthalpy. • We also

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consider entropy in the reaction. • Recall the first law of thermodynamics: energy is conserved. $\Delta E = q + w$

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Example 9.2 The element

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mercury, Hg, is a silvery liquid at room temperature. The normal freezing point of mercury is -38.9°C , and its molar enthalpy of fusion is $\Delta H_{\text{fusion}} = 2.29 \text{ kJ/mol}$. What is the entropy change of the
...

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Gravity. Created by.
Bryce_Naito PLUS. Terms in
this set (28) Enthalpy,
Entropy. What makes a
process, or more
specifically, a chemical
reaction, spontaneous? 1.
Tendency for a system to go

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Test Bank to a lower state of energy
(change in _____ will be a
...)

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(38) 1) The first law of thermodynamics can be given

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Test Bank. A) $\Delta E = q + w$ B) Δ
= - C) for any spontaneous
process, the entropy of the
universe increases

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Chemical Thermodynamics ΔS
For Phase Change Consider
the melting of ice • At 1
atm and 0 °C ice and water
are in equilibrium with each
other. • If we melt one mole
of ice at 0 °C/1 atm to form

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1 mole of water still at 0 °C • We need to give the water heat that would be equal to ΔH_{fusion} . • If the process is done extremely slowly the process

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- Chapter 19 Chemical Thermodynamics • Enthalpy •
A thermodynamic quantity that equal to the internal energy of a system plus the product of its volume and

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Entropy and the Second Law of Thermodynamics . Entropy Change • Entropy, S , is a thermodynamic term that reflects the disorder, or randomness, of the system. • The more disordered, or random, the system is, the

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Larger the value of S .

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thermodynamic quantity that
equal to the internal energy
of a system plus the product
of its volume and pressure
exerted on it by its

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 $H = E + pV$ Chemical Thermodynamics

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Chemical thermodynamics
(thermochemistry) is the
study of energy flow during
a chemical reaction. ...

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Test Bank There are 3 “laws of thermodynamics”. In Chapter 6, you learned the first law (conservation of energy: there is a constant amount of energy in the universe).
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I'll teach you how to
calculate the Gibbs Free
Energy Change (ΔG_f°) for

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(38) 1) The first law of thermodynamics can be given as _____. A) $\Delta E = q + w$ B) $\Delta = -$ C) for any spontaneous process, the entropy of the ...

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