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solutions. We evaluate the quantities of higher order derivative than the normal thermodynamic functions. We allow these higher derivative data speak for themselves without resorting to any model system. We thus elucidate the molecular processes in solution, (referred to in this book “mixing scheme ), to the depth equal to, if not deeper, than that gained by spectroscopic and other methods. We show that there are three composition regions in aqueous solutions of non-electrolytes, each of which has a qualitatively distinct mixing scheme. The boundary between the adjacent regions is associated with an anomaly in the third derivatives of  $G$ . The loci of the anomalies in the temperature-composition field form the line sometimes referred as “Koga line . We then take advantage of the anomaly of a third derivative quantity of 1-propanol in the ternary aqueous solution, 1-propanol – sample species –  $H_2O$ . We use its

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induced change as a probe of the effect of a sample species on H<sub>2</sub>O. In this way, we clarified what a hydrophobe, or a hydrophile, and in turn, an amphiphile, does to H<sub>2</sub>O. We also apply the same methodology to ions that have been ranked by the Hofmeister series. We show that the kosmotropes (salting out, or stabilizing agents) are either hydrophobes or hydration centres, and that chaotropes (salting in, or destabilizing agents) are hydrophiles. A new differential approach to solution thermodynamics A particularly clear elucidation of the mixing schemes in aqueous solutions A clear understandings on the effects of hydrophobes, hydrophiles, and amphiphiles to H<sub>2</sub>O A clear understandings on the effects of ions on H<sub>2</sub>O in relation to the Hofmeister effect A new differential approach to studies in multi-component aqueous solutions

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