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thermodynamics and  
use it for the study of  
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system. We thus elucidate the molecular processes in solution,

(referred to in this book “mixing scheme ), to the

depth equal to, if not

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by spectroscopic and

other methods. We

show that there are three

composition regions in

aqueous solutions of

non-electrolytes, each of

which has a

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qualitatively distinct mixing scheme. The boundary between the adjacent regions is associated with an anomaly in the third derivatives of  $G$ . The loci of the anomalies in the temperature-composition field form the line sometimes referred as “Koga line”. We then take advantage of the anomaly of a third

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derivative quantity of

1-propanol in the

ternary aqueous

solution, 1-propanol –

sample species – H<sub>2</sub>O.

We use its induced

change as a probe of the

effect of a sample

species on H<sub>2</sub>O. In this

way, we clarified what a

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hydrophile, and in turn,

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same methodology to ions that have been ranked by the Hofmeister series. We show that the kosmotropes (salting out, or stabilizing agents) are either hydrophobes or hydration centres, and that chaotropes (salting in, or destabilizing agents) are hydrophiles. A new differential

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