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$\text{CuO (s)} + \text{H}_2\text{(g)} \rightarrow \text{Cu (s)} + \text{H}_2\text{O (g)}$ Here, the oxidation number of Cu decreases from +2 in CuO to 0 in Cu i.e., CuO

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is reduced to
Cu. Also, the
oxidation number
of H increases
from 0 in H_2 to
+1 in H_2O
i.e., H_2 is
oxidized to H_2O .
Hence, this
reaction is a
redox reaction.

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primarily deals
in understanding
the classical

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idea of Redox
Reactions, that
involves the
oxidation and
reduction
reactions and
other topics
such as
electrode
processes,
oxidation
number, and
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as one of the
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TERMS OF
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TRANSFER

REACTIONS We

have already

learnt that the

reactions $2\text{Na}(s)$

$+ \text{Cl}_2(g) \rightarrow 2\text{NaCl}$

(s) (8.12)

$4\text{Na}(s) + \text{O}_2(g) \rightarrow$

$2\text{Na}_2\text{O}(s)$ (8.13)

$2\text{Na}(s) + \text{S}(s) \rightarrow$

$\text{Na}_2\text{S}(s)$ (8.14)

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are redox reactions because in each of these reactions sodium is oxidised due to the addition of either oxygen or more

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change in
oxidation states
of at least two

elements by
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involves two

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half-reactions,
one involving
loss of electron
or electrons
(oxidation) and
the other
involving gain
of electron or
electrons

(reduction). ION
ELECTRON METHOD
FOR BALANCING
REDOX REACTIONS.

This method

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involves the
following steps:

- (a) Divide the complete equations into two half reactions—(i) One representing oxidation

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Redox Reactions

Oxidation is
loss of

electrons and

reduction is

gain of

electrons.

Oxidation and

reduction always

occur

simultaneously,

thus the name

REDOX reaction.

Elements which

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lose electrons
are reducing
agents and those
which gain
electrons are
oxidising
agents.

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Reactions chapter
is an essential
chapter of the
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syllabus. The
chapter has high
importance in
the board exams
as well as in
the JEE and NEET
examinations.

So, the students
must have
adequate

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knowledge in
this chapter.
The important
topics and sub-
topics are given
below:

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...

In the
reaction, since

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H₂O₂ oxidise, O₃

into O₂ thus it
behaves as an

oxidising

agent.the

further

reaction, in the

reaction, Here

H₂O₂ reduces

Ag₂O into

metallic silver

[Ag] (as

oxidation number

is reducing from

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+1 to 0). Thus,
H₂O₂ behaves as
a reducing
agent.

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Carbon nanotubes are a recently discovered

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allotropes of
carbon. Three
other allotropes
of carbon—
buckyballs,
graphite, and
diamond—are
illustrated at
the left, as is
the molecule
methane, CH_4 ,
from which
nanotubes and
buckyballs can

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be made. The element carbon forms an amazing number of compounds with structures that follow from simple methane, found in natural gas, to the complex macromolecules that serve as the basis of

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sensing.
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proteins, there

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to develop drugs
that intercept
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pathways that
critically
involve nucleic
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naturally relies

on the

availability of

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methods that

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between nucleic

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tremendously
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both rapidly and
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