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Atkins' Physical Chemistry is always a recommended text; if you are starting a university chemistry course, I would highly recommend this book. All of the topics I have covered in university so far have been well explained by the textbook, with examples, diagrams and pictures where appropriate.

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A Textbook of Physical Chemistry, Second Edition serves as an introductory text to physical chemistry. Topics covered range from wave mechanics and chemical bonding to molecular spectroscopy and photochemistry; ideal and nonideal gases; the three laws of thermodynamics; thermochemistry; and solutions of nonelectrolytes. The kinetics of gas-phase reactions; colloids and macromolecules; and nuclear chemistry and radiochemistry are also discussed. This edition is comprised of 22 chapters; the first of which introduces the reader to the behavior of ideal and nonideal gases, with particular emphasis on the van der Waals equation. The discussion then turns to the kinetic molecular theory of gases and the application of the Boltzmann principle to the treatment of molar polarization; dipole and magnetic moments; the phenomenology of light absorption; and classical and statistical thermodynamics. The chapters that follow focus on the traditional sequence of chemical and phase equilibria, electrochemistry, and chemical kinetics in gas phase and solution phase. This book also considers wave mechanics and its applications; molecular spectroscopy and photochemistry; and the excited state, and then concludes with an analysis of crystal structure, colloid and polymer chemistry, and radio and nuclear chemistry. This reference material is intended primarily as an introductory text for students of physical chemistry.

A Textbook of Physical Chemistry: Second Edition provides both a traditional and theoretical approach in the study of physical chemistry. The book covers subjects usually covered in chemistry

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textbooks such as ideal and non-ideal gases, the kinetic molecular theory of gases and the distribution laws, and the additive physical properties of matter. Also covered are the three laws of thermodynamics, thermochemistry, chemical equilibrium, liquids and their simple phase equilibria, the solutions of nonelectrolytes, and heterogenous equilibrium. The text is recommended for college-level chemistry students, especially those who are in need of a textbook for the subject.

Written primarily to meet the requirements of students at the undergraduate level, this book aims for a self-learning approach. The fundamentals of physical chemistry have been explained with illustrations, diagrams, tables, experimental techniques and solved problems.

An advanced-level textbook of physical chemistry for the graduate (B.Sc) and postgraduate (M.Sc) students of Indian and foreign universities. This book is a part of four volume series, entitled "A Textbook of Physical Chemistry – Volume I, II, III, IV".

CONTENTS: Chapter 1. Quantum Mechanics – I: Postulates of quantum mechanics; Derivation of Schrodinger wave equation; Max-Born interpretation of wave functions; The Heisenberg's uncertainty principle; Quantum mechanical operators and their commutation relations; Hermitian operators (elementary ideas, quantum mechanical operator for linear momentum, angular momentum and energy as Hermitian operator); The average value of the square of Hermitian operators; Commuting operators and uncertainty principle(x & p ; E & t); Schrodinger wave equation for a particle in one dimensional box; Evaluation of average position, average momentum and determination of uncertainty in position and momentum and hence Heisenberg's uncertainty principle; Pictorial representation of the wave equation of a particle in one dimensional box and its influence on the kinetic energy of the particle in each successive quantum level; Lowest energy of the

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particle. Chapter 2. Thermodynamics – I: Brief resume of first and second Law of thermodynamics; Entropy changes in reversible and irreversible processes; Variation of entropy with temperature, pressure and volume; Entropy concept as a measure of unavailable energy and criteria for the spontaneity of reaction; Free energy, enthalpy functions and their significance, criteria for spontaneity of a process; Partial molar quantities (free energy, volume, heat concept); Gibb's-Duhem equation. Chapter 3. Chemical Dynamics – I: Effect of temperature on reaction rates; Rate law for opposing reactions of 1st order and 2nd order; Rate law for consecutive & parallel reactions of 1st order reactions; Collision theory of reaction rates and its limitations; Steric factor; Activated complex theory; Ionic reactions: single and double sphere models; Influence of solvent and ionic strength; The comparison of collision and activated complex theory. Chapter 4. Electrochemistry – I: Ion-Ion Interactions: The Debye-Huckel theory of ion-ion interactions; Potential and excess charge density as a function of distance from the central ion; Debye Huckel reciprocal length; Ionic cloud and its contribution to the total potential; Debye - Huckel limiting law of activity coefficients and its limitations; Ion-size effect on potential; Ion-size parameter and the theoretical mean-activity coefficient in the case of ionic clouds with finite-sized ions; Debye - Huckel-Onsager treatment for aqueous solutions and its limitations; Debye-Huckel-Onsager theory for non-aqueous solutions; The solvent effect on the mobility at infinite dilution; Equivalent conductivity (Λ) vs. concentration $c^{1/2}$ as a function of the solvent; Effect of ion association upon conductivity (Debye- Huckel - Bjerrum equation). Chapter 5. Quantum Mechanics – II: Schrodinger wave equation for a particle in a three dimensional box; The concept of degeneracy among energy levels for a particle in three dimensional box; Schrodinger wave equation for a linear harmonic oscillator & its solution by polynomial method; Zero point energy of a particle possessing harmonic motion and its consequence; Schrodinger wave equation for three dimensional Rigid rotator; Energy of rigid

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rotator; Space quantization; Schrodinger wave equation for hydrogen atom, separation of variable in polar spherical coordinates and its solution; Principle, azimuthal and magnetic quantum numbers and the magnitude of their values; Probability distribution function; Radial distribution function; Shape of atomic orbitals (s,p & d). Chapter 6. Thermodynamics – II: Classius-Clayperon equation; Law of mass action and its thermodynamic derivation; Third law of thermodynamics (Nernst heat theorem, determination of absolute entropy, unattainability of absolute zero) and its limitation; Phase diagram for two completely miscible components systems; Eutectic systems, Calculation of eutectic point; Systems forming solid compounds $A_x B_y$ with congruent and incongruent melting points; Phase diagram and thermodynamic treatment of solid solutions. Chapter 7. Chemical Dynamics – II: Chain reactions: hydrogen-bromine reaction, pyrolysis of acetaldehyde, decomposition of ethane; Photochemical reactions (hydrogen - bromine & hydrogen -chlorine reactions); General treatment of chain reactions (ortho-para hydrogen conversion and hydrogen - bromine reactions); Apparent activation energy of chain reactions, Chain length; Rice-Herzfeld mechanism of organic molecules decomposition(acetaldehyde); Branching chain reactions and explosions (H_2-O_2 reaction); Kinetics of (one intermediate) enzymatic reaction : Michaelis-Menton treatment; Evaluation of Michaelis 's constant for enzyme-substrate binding by Lineweaver-Burk plot and Eadie-Hofstae methods; Competitive and non-competitive inhibition. Chapter 8. Electrochemistry – II: Ion Transport in Solutions: Ionic movement under the influence of an electric field; Mobility of ions; Ionic drift velocity and its relation with current density; Einstein relation between the absolute mobility and diffusion coefficient; The Stokes- Einstein relation; The Nernst -Einstein equation; Walden 's rule; The Rate-process approach to ionic migration; The Rate process equation for equivalent conductivity; Total driving force for ionic transport, Nernst - Planck Flux equation; Ionic drift and diffusion potential; the Onsager

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phenomenological equations; The basic equation for the diffusion; Planck-Henderson equation for the diffusion potential.

Coverage of Physical Chemistry. Each volume includes a large number of illustrative numericals and typical problems to highlight the principles involved. IUPAC recommendations and SI units have been adopted throughout. The present book describes Wave Mechanics, Energy Quantization and Atomic Structure, Theories of Covalent Bond, Electrical and Magnetic Properties of Molecules, Molecular Spectroscopy, Molecular Symmetry and its Applications. Salient Features: • Comprehensive coverage of wave mechanics, energy quantization and atomic structure, theories of covalent bond, electrical and magnetic properties of molecules, molecular spectroscopy, molecular symmetry and its applications • Emphasis given to applications and principles • Explanation of equations in the form of solved problems and numericals • IUPAC recommendations and SI units have been adopted throughout • Rich and illustrious pedagogy

This elegant book provides a student-friendly introduction to the subject of physical chemistry. It is concise and more compact than standard textbooks on the subject and it emphasises the two important concepts underpinning physical chemistry: quantum mechanics and the second law of thermodynamics. The principles are challenging to students because they both focus on uncertainty and probability. The book explains these fundamental concepts clearly and shows how they offer the key to understanding the wide range of chemical phenomena including atomic and molecular spectra, the structure and properties of solids, liquids and gases, chemical equilibrium, and the rates of chemical reactions.

Originally published in 1950, this textbook was intended for school students with the aim of providing an introductory understanding of chemistry. The book introduces physical chemistry through multiple

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and diverse experiments; each experiment designed to reinforce a new topic and reflect theorems, approaches and historical development. Notably, the treatment throughout is from the point of view of the kinetic-molecular theory rather than that of the laws of thermodynamics, whilst emphasis is also placed upon physico-chemical phenomena and their significance in various branches of science, such as metallurgy, chemical syntheses and mineralogy. There are twelve chapters in total, with chapter titles ranging from 'Atoms and molecules' to 'Mass action and the ionic dissociation theory'. Various diagrams and plate sections are also included for reference. This book will be of value to chemistry students and scholars as well as those interested in the history of education.

This comprehensive textbook, now in its second edition, is mainly written as per the latest syllabi of physical chemistry of all the leading universities of India as well as the new syllabus recommended by the UGC. This thoroughly revised and updated edition covers the principal areas of physical chemistry, such as thermodynamics, quantum chemistry, molecular spectroscopy, chemical kinetics, electrochemistry and nanotechnology. In a methodical and accessible style, the book discusses classical, irreversible and statistical thermodynamics and statistical mechanics, and describes macroscopic chemical systems, steady states and thermodynamics at a molecular level. It elaborates the underlying principles of quantum mechanics, molecular spectroscopy, X-ray crystallography and solid state chemistry along with their applications. The book explains various instrumentation techniques such as potentiometry, polarography, voltammetry, conductometry and coulometry. It also describes kinetics, rate laws and chemical processes at the electrodes. In addition, the text deals with chemistry of corrosion and nanomaterials. This text is primarily designed for the undergraduate and postgraduate students of chemistry (B.Sc. and M.Sc.) for their course in physical chemistry. Key Features • Gives a thorough treatment to ensure a

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solid grasp of the material. • Presents a large number of figures and diagrams that help amplify key concepts. • Contains several worked-out examples for better understanding of the subject matter. • Provides numerous chapter-end exercises to foster conceptual understanding.

Emphasizes a molecular approach to physical chemistry, discussing principles of quantum mechanics first and then using those ideas in development of thermodynamics and kinetics. Chapters on quantum subjects are interspersed with ten math chapters reviewing mathematical topics used in subsequent chapters. Includes material on current physical chemical research, with chapters on computational quantum chemistry, group theory, NMR spectroscopy, and lasers. Units and symbols used in the text follow IUPAC recommendations. Includes exercises. Annotation copyrighted by Book News, Inc., Portland, OR

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